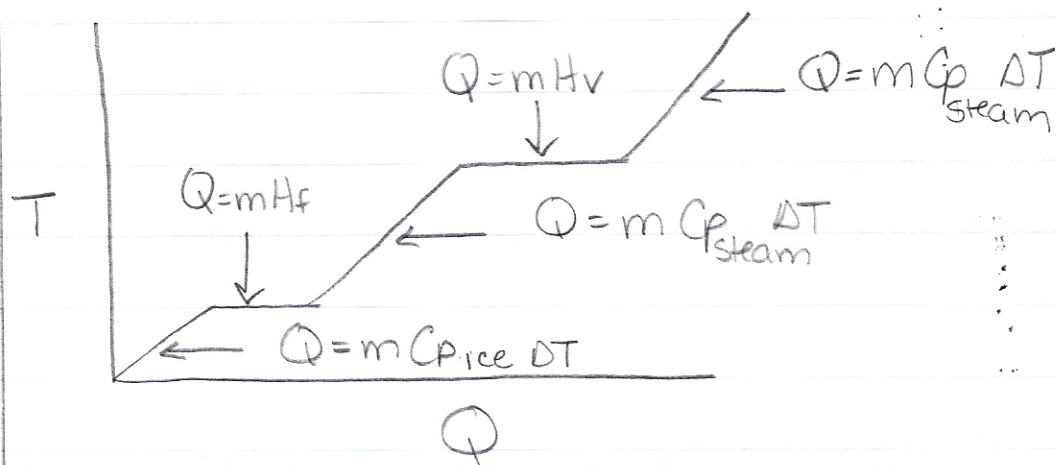


molar
heat of
vaporization

Energy required to change 1 mol
liquid into vapor
 $= 40.6 \frac{\text{kJ}}{\text{mol}}$ for H_2O

equations



H_f

Heat of fusion = 80.0 cal/g
 334 J/g

H_v

Heat of vaporization = $540. \text{ cal/g}$
 2260 J/g

$$C_{p,steam} = 2.06 \text{ J/g}^\circ\text{C}$$

$$C_{p,ice} = 2.03 \text{ J/g}^\circ\text{C}$$

$$C_{p,H_2O} = 4.184 \text{ J/g}^\circ\text{C} = 1.00 \text{ cal/g}^\circ\text{C}$$